

(12) INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(19) World Intellectual Property Organization
International Bureau(43) International Publication Date
9 October 2003 (09.10.2003)

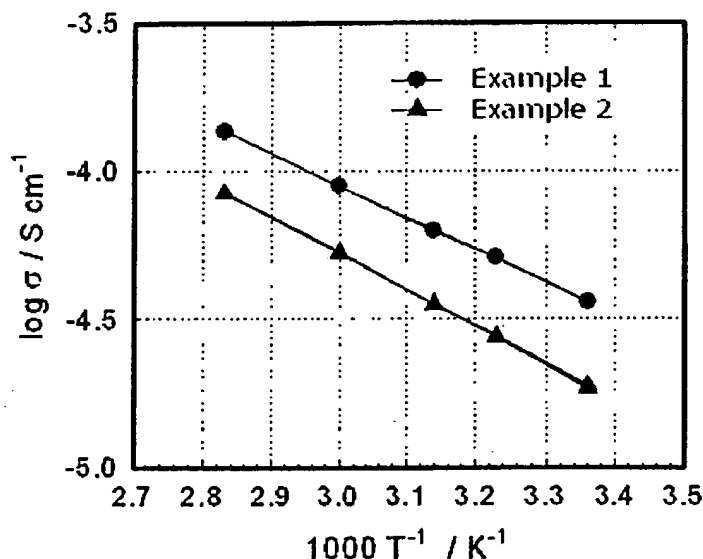
PCT

(10) International Publication Number
WO 03/083971 A1

- (51) International Patent Classification⁷: **H01M 6/14**,
6/18, 10/40, H01G 9/025
- (21) International Application Number: PCT/US03/02128
- (22) International Filing Date: 22 January 2003 (22.01.2003)
- (25) Filing Language: English
- (26) Publication Language: English
- (30) Priority Data:
10/104,352 22 March 2002 (22.03.2002) US
10/167,940 12 June 2002 (12.06.2002) US
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- (81) Designated States (national): AE, AG, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, BZ, CA, CH, CN, CO, CR, CU, CZ, DE, DK, DM, DZ, EC, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MA, MD, MG, MK, MN, MW, MX, MZ, NO, NZ, OM, PH, PL, PT, RO, RU, SC, SD, SE, SG, SK, SL, TJ, TM, TN, TR, TT, TZ, UA, UG, US, UZ, VC, VN, YU, ZA, ZM, ZW.
- (84) Designated States (regional): ARIPO patent (GH, GM, KE, LS, MW, MZ, SD, SL, SZ, TZ, UG, ZM, ZW), Eurasian patent (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European patent (AT, BE, BG, CH, CY, CZ, DE, DK, EE, ES, FI, FR, GB, GR, HU, IE, IT, LU, MC, NL, PT, SE, SI, SK, TR), OAPI patent (BF, BJ, CF, CG, CI, CM, GA, GN, GQ, GW, ML, MR, NE, SN, TD, TG).

[Continued on next page]

(54) Title: SOLID POLYMER ELECTROLYTE AND METHOD OF PREPARATION



(57) Abstract: Disclosed is an improved solid electrolyte made of an interpenetrating network type solid polymer comprised of two compatible phases: a crosslinked polymer for mechanical strength and chemical stability, and an ionic conducting phase. The highly branched siloxane polymer of the present invention has one or more poly(ethylene oxide) ("PEO") groups as a side chain. The PEO group is directly grafted to silicon atoms in the siloxane polymer. This kind of branched type siloxane polymer is stably anchored in the network structure and provides continuous conducting paths in all directions throughout the IPN solid polymer electrolyte. Also disclosed is a method of making an electrochemical cell incorporating the electrolyte. A cell made accordingly has an extremely high cycle life and electrochemical stability.



Published:

- with international search report
- with amended claims

For two-letter codes and other abbreviations, refer to the "Guidance Notes on Codes and Abbreviations" appearing at the beginning of each regular issue of the PCT Gazette.

Solid Polymer Electrolyte and Method of Preparation

CONTRACTUAL ORIGIN OF THE INVENTION

[0001] This invention was made with United States Government support under NIST
5 ATP Award No. 70NANB043022 awarded by the National Institute of Standards and
Technology (NIST). The United States Government has certain rights in this invention
pursuant to NIST ATP Award No. 70NANB043022 and pursuant to Contract No. W-31-109-
ENG-38 between the United States Government and the University of Chicago representing
Argonne National Laboratory.

10 FIELD

[0002] The present invention relates to the composition and assembly methods of solid
polymer electrolytes and their use in electrochemical cells, especially in lithium ion
rechargeable batteries. The invention particularly relates to interpenetrating network type
solid polymer electrolyte systems with highly ionic conductivity poly(ethylene oxide)
15 ("PEO") grafted siloxane polymers as a conducting phase.

BACKGROUND

[0003] Efforts to develop electrochemical cells having PEO based solid electrolyte
systems have continued since about 1973. (M.B. Armand, Fast Ion Transport in Solids,
North Holland, Amsterdam, p665, (1973); D.E. Fenton et al., Polymer, 14, 589 (1973)). The
20 main advantages of such a cell system are multifold: (1) very high energy density;
(2) potential for excellent electrolyte stability; (3) the ability to be configured in nearly any
shape since it contains no liquid; (4) the opportunity to be very inexpensive; (5) inherent
safety characteristics; and (6) an expansive market if successfully developed. Up to now the
key impediment to the successful development of such a polymer cell for room temperature
25 operation is the low ionic conductivity of the solid polymer electrolyte. A major effort to
develop the solid polymer electrolyte ("SPE") system is being carried out by Hydro Quebec
and 3M under contract to the United States Advanced Battery Consortium (USABC) for
electric vehicle applications. The batteries developed in this effort are operated at
approximately 60°C to 80°C (140°F to 176°F), and achieve about 800 cycles (M. Gauthier et
30 al., J. Power Sources, 54, 163 (1995)). All attempts in this program to successfully develop a
room temperature SPE based battery were unsuccessful because of the low ionic conductivity
at room temperature of PEO based electrolyte using the lithium trifluoromethane sulfonyl

imide [LiN(CF₃SO₃), LiTFSI] salt ("TFSI"). Based on examination and evaluation of the various solid electrolytes developed to date (L.A. Dominey et al., *Electrochim. Acta*, 37, 1551 (1992); F. Alloin et al., *Solid State Ionics*, 60, 3 (1993)), it is quite apparent that PEO based polymer or derivative thereof appears to be the most promising.

- 5 [0004] One type of PEO investigated thoroughly is the high molecular weight (about 4 million) linear variety, which forms relatively strong, free-standing films at room temperature. Its strength is derived from a semicrystalline microstructure. Lithium ion transport in such materials depends on the complexation of lithium ions by the oxygen atoms in oxyethylene units in the polymer chains. High molecular weight PEO doped with the lithium salt
- 10 LiN(SO₂CF₃)₂, LiTFSI, has an optimum conductivity of 10⁻⁵ S/cm at 80°C (176°F) (S. Kohama et al., *J. Appl. Polym. Sci.*, 21, 863 (1977)). Many lithium salt complexes of PEO at room temperature are predominantly crystalline until a melting point of 68°C (154.5°F) leading to very poor ionic conductivities of approximately 10⁻⁷ S/cm. The improved conductivities using the TFSI salt are due to the plasticizing effect of the anion
- 15 which substantially reduces the crystallinity of the PEO complex at room temperature. It is important to note that only the amorphous PEO electrolyte is ionically conductive.

- [0005] Substantial research effort has been devoted to lowering the operating temperature of SPE to the ambient region. To solve this problem, alkyl phthalates and poly(ethylene glycol) dialkyl ether with low molecular weight have been used as plasticizing additives for
- 20 SPE to reduce the crystalline region and increase the mobility of the SPE molecular chain at ambient temperature. Low molecular poly(ethylene oxide - dialkyl ether compounds) can contribute to increased room temperature ionic conductivity of SPE, but they still have a crystallization problem which decreases the ionic conductivity at certain temperatures. Another approach to attempt to improve the ionic conductivity at ambient temperature was to
- 25 synthesize a highly branched PEO to decrease the crystalline tendency of PEO main chain and to increase the chain mobility regarding lithium ion transport such as hyper-branched SPE (Z. Wang et al., *J. Electrochem. Soc.*, 146(6), 2209 (1999)), and comb-like SPEs (J.S. Gnanaraj, R.N. Karekar et al., *Polymer*, 38(14) 3709 (1997)). However, the ionic conductivity of such highly branched PEO is still low at ambient temperature. All of these
- 30 efforts were intended to create amorphous polymer near ambient temperature.

[0006] To apply a SPE to a real practical electrochemical cell system, adequate mechanical strength is required. Simple crystalline PEO may meet that requirement, but most of modified PEO based SPEs are not strong enough for real cell applications.

Crosslinked SPEs were developed as a solution, but the crosslinking reaction restricts polymer chain mobility that is needed for lithium ion transport. (U.S. Pat. Nos. 4,908,283 to Takahashi, 4,830,939 to Lee, 5,037,712 to Shackle and 3,734,876 to Chu). More advanced systems are the interpenetrating network ("IPN") type SPE that consist of crosslinked polymers and an ionic conducting phase which is mostly low molecular weight PEO base compounds. The ionic conductivity of such systems, however, still depends on the flexibility of poly(alkylene oxide) which has a temperature dependency on its mobility, as well as on its mechanical strength which is not obvious. Most prior patents have disclosed the use of volatile solvents to dissolve the PEO compounds and metal salts. The use of volatile solvents to make the SPEs increase the processing steps such as evaporation and recovery, increase costs of manufacture, and may pose serious environmental and safety issues. U.S. Patent No. 5,112,512 to Nakamura discloses crosslinking PEO crosslinking agent to siloxane with a PEO side chain which has a reactive unsaturated bond. This crosslinking approach results in a significantly reduced flexibility of siloxane with PEO polymer. The present invention is distinguished in that the siloxane is captured inside the network with no chemical bonds to the PEO crosslinking agent, greatly enhancing flexibility.

[0007] Accordingly, the present inventors have developed a new type of IPN polymer electrolyte having PEO grafted onto polysiloxanes as an ion conducting phase and a porous support to overcome the above-mentioned problems such as low room temperature ionic conductivity, chemical and electrochemical stability, as well as safety. The PEO grafted polysiloxanes are liquid compounds, electrochemically stable and have low glass transition temperature with little or no crystallization problems. Notably the present invention does not include any volatile solvent in the polymer electrolyte preparation.

OBJECTIVES

[0008] A primary objective of the present invention is to provide an IPN SPE having increased room temperature ionic conductivity with chemical and electrochemical stability.

[0009] Another object of the invention is to provide a thin IPN SPE with reduced bulk impedance and excellent mechanical strength.

[0010] An additional object of the present invention is to provide an improved method to manufacture an electrochemical battery having an IPN solid polymer.

[0011] To fulfill the above objectives, the IPN SPE in the present invention is fabricated by using the composition which comprises branched type siloxane polymer in a liquid state, crosslinking agent selected from diacrylate terminated poly(alkylene oxide) compounds, crosslinking density controlling compounds selected from poly(alkylene oxide) acrylate alkyl ether compounds, a lithium salt and a thermal initiator.

[0012] A further object of the invention is to provide a fabrication method to prepare the IPN type SPEs through thermal crosslinking. This method uses a porous media such as polyolefin separator, nonwoven fabrics, polycarbonate membrane, etc. to reduce the bulk impedance of SPE through minimizing its thickness.

10 SUMMARY

[0013] The present invention relates to a SPE and its preparation method. More particularly, the present invention relates to an IPN type SPE having PEO grafted siloxane polymer as the major ionic conducting phase which has excellent ionic conductivity and mechanical strength, and its preparative method is quite simple. The SPEs resulting from the present invention are suitable electrolytes for lithium polymer secondary batteries.

BRIEF DESCRIPTION OF THE DRAWINGS

[0014] Fig. 1 is a temperature vs. ionic conductivity graph for the present invention.

[0015] Fig. 2 is a flow chart illustrating the steps of a method for implementing the present invention.

20 [0016] Fig. 3 is a flow chart illustrating the steps of another method for implementing the present invention.

[0017] Fig. 4 is a flow chart illustrating the steps of another method for implementing the present invention.

[0018] Fig. 5 is a flow chart illustrating the steps of another method for implementing the present invention.

[0019] Fig. 6 is a graph showing the cyclic voltammogram of electrochemical stability.

[0020] Figs. 7a, 7b and 7c are graphs showing the measured heat flow from decomposition reaction of test samples.

[0021] Figs. 8a and 8b show a potentiostatic curve (Fig. 8a) and impedance spectra (Fig. 8b) of lithium metal/IPN SPE of Example 2/lithium metal cell to measure lithium transference number of the SPE.

5 [0022] Figs. 9a and 9b show a charge/discharge pattern (Fig. 9a) and specific discharge capacity (Fig. 9b) according to cycle number of lithium metal/IPN SPE of Example 8/LiNi_{0.8}Co_{0.2}O₂ cathode cell.

[0023] Fig. 10a is an illustration of the method of fabricating a cell according to the present invention.

[0024] Fig. 10b is a cut-away drawing of an electrochemical cell incorporating a SPE.

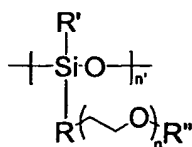
10 DETAILED DESCRIPTION

[0025] The SPE of the present invention comprises an IPN of two separate continuous phases that are compatible with each other. One of the phases is a crosslinked polymer that ensures its mechanical strength and chemical stability, and the other is a conducting phase for dissociating ion. The crosslinking phase can also assist metal salt dissolution and
15 transportation.

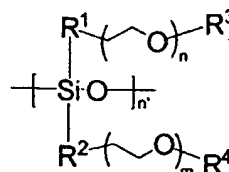
[0026] The elaborately designed highly branched siloxane polymer of the present invention has one or more PEO groups as a side chain. The PEO group is directly grafted to silicon atoms in the siloxane polymer. This kind of branched type siloxane polymer is stably anchored in the network structure and provides continuous conducting paths in all directions
20 throughout the IPN SPE.

[0027] The branched type siloxane polymer easily dissolves the lithium salt and has the required flexibility to transport the lithium ions. Through the fabrication method suggested by this invention, the polysiloxane is well anchored in the IPN polymer electrolyte and increases the polymer ionic conductivity by its high segmental mobility.

25 [0028] The present invention includes all types of siloxane polymers with PEO as a side chain and the branched type siloxane polymers represented by the formula (I-a and I-b) are specific examples of the present invention:



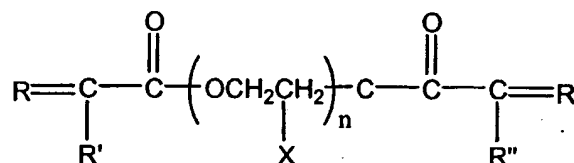
(I-a)



(I-b)

[0029] Wherein each of R, R¹ and R² represents oxygen or a group selected from an
alkylene oxide group having 1 to 6 carbon atoms; each of R', R'', R³ and R⁴ represents
5 hydrogen or a group selected from an alkyl group having 1 to 12 carbon atoms and/or an
alkenyl group having 2 to 8 carbon atoms; each of n and m represents whole numbers from 1
to 12; and n' represents whole numbers from 4 to 10,000.

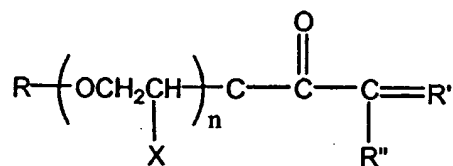
[0030] The crosslinking agent in the present invention is represented by formula (II):



(II)

[0031] Wherein R represents a group selected from an alkyl group having 1 to 10 carbon
atoms; and each of R' and R'' represents hydrogen or a group selected from an alkyl group
having 1 to 10 carbon atoms and/or an alkenyl group having 2 to 12 carbon atoms; and X
being hydrogen or a methyl group; and n represents a numeral of 1 to 15.

15 [0032] The monomer used for the control of crosslinking density of the IPN SPE is
represented by formula (III):



(III)

[0033] Wherein each of R and R' represents a group selected from an alkyl group having
20 1 to 10 carbon atoms; and R'' represents hydrogen or a group selected from an alkyl group
having 1 to 10 carbon atoms and/or an alkenyl group having 2 to 12 carbon atoms; and X is
hydrogen or a methyl group; and n represents a numeral of 1 to 20.

[0034] The lithium salt to be used in the present invention is not particularly limited, as long as it dissolves in the polymer and serves as an electrolyte for a lithium secondary battery. Examples of specific lithium salts include LiClO_4 , LiBF_4 , LiAsF_6 , LiPF_6 , LiCF_3SO_3 , $\text{Li}(\text{CF}_3\text{SO}_2)_2\text{N}$, $\text{Li}(\text{CF}_3\text{SO}_2)_3\text{C}$, $\text{LiN}(\text{SO}_2\text{C}_2\text{F}_5)_2$, lithium alkyl fluorophosphates, lithium bis(oxalato) borate (LiBOB), and a mixture thereof, as well as salts yet to be identified.

[0035] The molar ratio of the lithium salt to the oxygen in the organic mixture of branched type siloxane polymer, crosslinking agent and monofunctional compound is preferably 0.01 to 0.2. If the proportion of the lithium salt is smaller than 0.01, the ionic conductivity of the resulting IPN SPE is significantly decreased because of an inadequate number of carrier ions are in the SPE. If the molar ratio is greater than 0.2, the lithium salt is not sufficiently dissociated in the resulting IPN SPE and the aggregation of lithium ion can reduce the ionic conductivity.

[0036] Figure 1 shows the effect of temperature on the ionic conductivity of two IPN polymer electrolytes. Example 1 is with 60wt% branched type siloxane polymer ($n=7.2$ in formula I-a), 30wt% poly(ethylene glycol) ethyl ether methacrylate, 10wt% of poly(ethylene glycol-600) dimethacrylate and $\text{Li}(\text{CF}_3\text{SO}_2)_2\text{N}$. Example 2 is with 50wt% branched type siloxane polymer ($n=7.2$, R' and R'' are methyl groups in formula I-a), 40wt% poly(ethylene glycol) ethyl ether methacrylate, 10wt% of poly(ethylene glycol-600) dimethacrylate and $\text{Li}(\text{CF}_3\text{SO}_2)_2\text{N}$. Both of the two IPN SPEs show high ionic conductivity over 10^{-5} S/cm at room temperature. The anchored siloxane, without any chemical bonding with the IPN structure, gives improved ionic conductivity to the IPN type crosslinked SPE. More important is the content of liquid siloxane in the IPN SPE; as the more content of branched type siloxane polymer is increased, so is the ionic conductivity.

[0037] Figure 2 delineates the requisite steps in a method for preparing an interpenetrating network type polymer electrolyte, comprising the steps of: (1) dissolving a lithium salt and a thermal initiator selected from azo compounds such as azoisobutyronitrile, peroxide compounds such as benzoylperoxide and bismaleimide in a branched type siloxane based polymer (formula I-a or I-b); (2) mixing two or more acrylate terminated polyalkylether comprising either or both of an ethylene oxide unit and a propylene oxide unit as a monomeric unit in the resulting solution; (3) casting the resulting mixture called a precursor solution onto a substrate, porous medium such as polyolefin separator, nonwoven and polycarbonate membrane or a surface of the electrode; and (4) placing the cast film in an oven or on a heating medium for solidification thereof.

[0038] The porous media of the present invention will be used to reduce the thickness of IPN SPE and are preferably polyolefin separator, nonwovens and polycarbonate microporous membrane. The final thickness of SPE of the present invention with the porous supporter is below 100 μ m, preferably below 50 μ m.

5 [0039] Figure 3 delineates the method for assembling a lithium rechargeable cell with the SPE of the present invention, comprises the steps of: (1) coating the precursor solution including siloxane polymer, crosslinking agent, chain length controlling agent, lithium salt and radical initiator onto one or more surfaces coming from porous supporter, cathode laminate and anode laminate; (2) curing the precursor solution to make the SPE; (3) stacking
10 each components including porous supporter, cathode laminate and anode laminate properly; (4) winding or folding the stacked components to prepare the spiral wound cell type or prismatic cell type; and (5) packaging the cell in a metal can, plastic pouch or laminated plastic/metal foil pouch. Such stacking, winding and packaging are well known in the art.

[0040] Figure 4 denotes another method for assembling lithium rechargeable cell with the
15 SPE of the present invention, comprises the steps of: (1) coating the precursor solution including siloxane polymer, crosslinking agent, chain length controlling agent, lithium salt and radical initiator onto one or more surfaces coming from porous supporter, cathode laminate and anode laminate; (2) stacking each component, including porous supporter, cathode laminate and anode laminate properly; (3) winding or folding the stacked
20 components to prepare spiral wound cell type or prismatic cell type; (4) curing the cell to change the precursor solution into SPE; and (5) packaging the cell with metal can, plastic pouch or laminated foil/ plastic pouch.

[0041] Figure 5 denotes the steps of another method for assembling the lithium rechargeable cell with the SPE of the present invention, comprises the steps of: (1) stacking
25 each component including porous supporter, cathode laminate and anode laminate properly to assemble the cell; (2) winding or folding the stacked components to prepare spiral wound cell type or prismatic cell type; (3) putting the cell in a metal can, plastic pouch or laminated metal foil/plastic pouch; (4) injecting the precursor solution including siloxane polymer, crosslinking agent, chain length controlling agent, lithium salt and radical initiator into the
30 pre-assembled cell; and (5) curing the cell to change the precursor solution to SPE.

[0042] The present invention will be better understood by reference to the following examples which are intended for purposes of illustration and are not intended to nor are to be

interpreted in any way as limiting the scope of the present invention, which is defined in the claims appended hereto.

[0043] EXAMPLES 1-2

[0044] For Examples 1-2, $\text{Li}(\text{CF}_3\text{SO}_2)_2\text{N}$ (LiTFSI) salt was dissolved in a branched type siloxane polymer ($n=7.2$, R' and R'' are methyl groups in formula I-a, $\text{Mn}=\text{ca. } 2000$), poly(ethylene glycol) ethyl ether methacrylate (PEGEEMA) with average Mn of ca. 246 and poly(ethylene glycol-600) dimethacrylate (PEGDMA600) with average Mn of ca. 740 mixture. After clear dissolution of LiTFSI, benzoyl peroxide was added into the resulting solution and mixed to get a precursor solution for IPN type SPE. The composition of Examples 1-2 is shown in Table 1. Porous polycarbonate membrane is used as a supporter for the IPN SPEs. The ionic conductivity of the IPN polymer electrolytes at temperatures ranging from 25 to 80°C were measured from the ac impedance curves of 2030 button cells assembled by sandwiching the IPN SPE between two stainless steel discs with a frequency range from 1 MHz to 10 Hz. The result is shown in Fig. 1. Both of two IPN SPEs show high ionic conductivity over 10^{-5} S/cm at room temperature and as the content of branched type siloxane polymer is increased, so is the ionic conductivity.

Table. 1

[grams]

Example #	Composition				
	Branched siloxane polymer	PEGDMA600	PEGEEMA	LiTFSI	BPO
1	2.400	0.400	1.200	0.770	0.016
2	2.000	0.400	1.600	0.800	0.020

[0045] EXAMPLE 3

[0046] Figure 6 is a trace of current density vs. potential during repeated voltage sweep of a cell made according to this invention at a scan rate of 5mV/sec. It shows the electrochemical stability with 60wt% branched type siloxane polymer ($n=7.2$, R' and R'' are methyl groups in formula I-a infra), 30wt% poly(ethylene glycol) ethyl ether methacrylate, 10wt% of poly(ethylene glycol-600) dimethacrylate and $\text{Li}(\text{CF}_3\text{SO}_2)_2\text{N}$. More specifically, it shows the electrochemical stability of IPN SPE with the same composition as Example 1. Polypropylene melt-blown type nonwoven separator material is used as a supporter for this IPN SPE. The electrochemical stability window of this IPN polymer electrolyte was determined by cyclic voltammetry with a 2030 button cell assembled by sandwiching this

IPN SPE between a stainless steel disc as a working electrode and lithium metal disc as the counter and reference electrodes. This IPN SPE shows an excellent electrochemical stability window of over 4.5 V and only a minimal decomposition peak around 4.5 V during the first anodic sweep. Notably, except for a slight variation during the first cycle, each subsequent cycle shows almost identical current density versus potential. This level of electrochemical stability during repeated cycling is extraordinary.

[0047] COMPARISONS 1 AND 2, AND EXAMPLE 4

[0048] Accelerating rate calorimetry (ARC) was used to investigate the chemical and thermal degradation of branched type siloxane polymer and its IPN polymer electrolyte at elevated temperatures of up to 400°C. The ARC is an adiabatic calorimeter in which heat evolved from the test sample is used to raise the sample temperature. The ARC is conducted by placing a sample in a sample bomb inside an insulating jacket. In an ARC analysis, the sample is heated to a preselected initial temperature and held a period of time to achieve thermal equilibrium. A search is then conducted to measure the rate of heat gain (self-heating) of the sample. If the rate of self-heating is less than a preset rate after the programmed time interval (typically 0.02°C/min), the sample temperature is stepped to a new value, and the heat-wait-search sequence is repeated. Once a self-heating rate greater than the present value is measured, the heat-wait-search sequence is abandoned; the only heating supplied to the calorimeter thereafter is that required to maintain the adiabatic condition between the sample and the jacket. Heat generated from the reaction inside the sample increases its temperature and pressure, thereby increasing the rate of the reaction. Sample weight for the test was 500 mg. Each sample was introduced in a 2¼" x ¼" diameter stainless steel bomb as a sample for ARC test. The detail compositions are explained in Table 2.

Table 2

Sample	Branched siloxane polymer	PEGDMA600	[grams]
			LiTFSI
Comparison 1	0.5000	-	-
Comparison 2	2.0000	-	0.3205
Example 4	2.0000	1.3333	0.3495

[0049] Figures 7a-c show the heat flow from decomposition reaction of the samples. Figure 7a (Comparison 1) shows its rapid reaction due to thermal decomposition and carbonization around 350°C. (662°F). Figure 7b (Comparison 2) shows better thermal

behavior caused by the coordination bonds between oxygen atoms and the lithium salt, which has a thermal stability of over 350°C (662°F), but there was still decomposition reaction around 360-370°C (648°F-698°F). Figure 7c (Example 4) shows only the heat of reaction due to thermal crosslinking to make an IPN structure at 60°C (140°F) and then no significant decomposition reaction was detected up to 400°C (752°F). It was thus found that the IPN structure significantly enhances the thermal stability of its SPE.

[0050] EXAMPLE 7 – Lithium ion Transference Number

[0051] Li metal/IPN SPE of Example 2/Li metal cell was assembled for the measurement of Li ion transference number t_+ . A 2030 button cell was used. A potentiostatic curve (Fig. 8a) was measured by using dc polarization method and the change of the cell impedance before and after polarization (Fig. 8b) was examined by using Schlumberger model 1255 frequency response analyzer connected to Schlumberger model 1286 electrochemical interface and EG&G PAR 273 potentiostat. The Li transference number was given by following equation suggested by K.M. Abraham et al., Chem. Mater., 9, 1978 (1997):

$$t_+ = \frac{I_s R_b^s (V - I_o R_i^o)}{I_o R_b^o (V - I_s R_i^s)}$$

[0052] Wherein V is the dc potential applied across the symmetric cell, o and s represent the initial and steady state, and b and i represent bulk and interfacial resistance of the electrolyte.

[0053] Lithium transference number of IPN SPE of Example 2 was approximately 0.29, which is much improved over that of pure PEO SPE, which is about 0.015.

[0054] EXAMPLE 8 – Cell Test

[0055] IPN SPE was prepared using the same composition as in Example 1 with a nonwoven support material. A 2030 button cell was assembled with lithium metal as an anode, IPN SPE of Example 8 and $\text{LiNi}_{0.8}\text{Co}_{0.2}\text{O}_2$ as a cathode. The preparation method for assembling the 2030 button cell with the SPE of Example 8, comprised the steps of: coating the precursor solution of Example 8 onto cathode laminate; stacking IPN SPE and lithium metal; putting a plate spring and top lid on the stacked components to 2030 button cell and crimping; curing the cell to change the precursor solution to SPE at 70°C for 1 hr.

[0056] The composition of the cathode is listed in Table 3. The effective cell area was 1.6 cm^2 . Charge and discharge rate were C/6. There was no degradation peak caused by the metal oxide up to 4.1 V and the specific discharge capacity was over 130 mAh/g.

Table 3

Electrode	Composition [wt%]			
	$\text{LiNi}_{0.8}\text{Co}_{0.2}\text{O}_2$	PVdF*	Graphite	Carbon black
Cathode	84	8	4	4

5 * poly(vinylidene fluoride)

[0057] The remarkable electrochemical stability of the present invention is illustrated by Figures 9a and 9b. It can be seen in Figure 9a that after 48 cycles (charges and discharges) there was no decrease in capacity at all. Although the cathode material, $\text{LiNi}_{0.8}\text{Co}_{0.2}\text{O}_2$, is a strong oxidizing material, the IPN SPE of this invention does not show any degradation
 10 problem up to 4.1 V. Figure 9b shows that the capacity of the test cell made according to the present invention increased in capacity from 100 mAh/g to about 130 mAh/g over the first approximately 40 cycles, then remained stable at that level.

[0058] Figures 10a and 10b illustrate the construction of an electrochemical cell 200 incorporating the SPE of the best mode of the present invention. This prismatic "wound" type cell ("jelly roll") comprises (1) a cast positive electrode material 204 made of metal
 15 oxide active material, PVDF binder and carbon additive with the precursor solution made of the mixture of comb siloxane polymer, poly(ethylene glycol) ethyl ether methacrylate, poly(ethylene glycol) dimethacrylate and lithium salt $\text{Li}(\text{CF}_3\text{SO}_2)_2\text{N}$ 206, (2) porous polycarbonate film 208 cast with the precursor solution of comb siloxane polymer,
 20 poly(ethylene glycol) ethyl ether methacrylate, poly(ethylene glycol) dimethacrylate and lithium salt $\text{Li}(\text{CF}_3\text{SO}_2)_2\text{N}$ 207, and (3) a cast porous negative electrode 212 made of carbon and PVDF binder with the precursor solution made of the mixture of comb siloxane polymer, poly(ethylene glycol) ethyl ether methacrylate, poly(ethylene glycol) dimethacrylate and lithium salt $\text{Li}(\text{CF}_3\text{SO}_2)_2\text{N}$ 214. The positive and negative electrodes 204 and 212 are
 25 electrically coupled to terminals 216 and 220.

[0059] This invention is equally applicable to related technologies including super capacitors and hybrid devices incorporating aspects of capacitors and batteries. For the purposes of this patent, "electrochemical cell" shall refer to all forms of electrochemical

storage devices, including single cells, batteries, capacitors, super capacitors and hybrid electrochemical devices.

[0060] BEST MODE OF THE PRESENT INVENTION

[0061] The inventors believe the best mode for the IPN polymer electrolyte is with the composition of 10wt% to 80wt%, more preferably about 30wt% to 75wt%, even more preferably about 50wt% to 70wt%, comb type siloxane polymer ($n=7.2$ in formula I-a), 30wt% poly(ethylene glycol) ethyl ether methacrylate, 10wt% of poly(ethylene glycol) dimethacrylate with molecular weight of 600 and $\text{Li}(\text{CF}_3\text{SO}_2)_2\text{N}$ supported by porous polycarbonate membrane. This composition and construction shows the highest ionic conductivity of 3.6×10^{-5} S/cm at 25°C and 5.1×10^{-5} S/cm at 37°C. The crosslinking agent should constitute between 5% and 60%, more preferably between 10% and 40%, by weight of all organic compounds in the SPE. The monomeric compound for controlling crosslinking density should constitute about 15% to 40% by weight of the total weight of organic compounds in the SPE. The thickness of porous polycarbonate membrane should be approximately 20 μm and the total thickness of the IPN polymer electrolyte should be about 95 μm . The radical reaction initiator should be a thermal initiator such as benzoyl peroxide and azoisobutyronitrile.

[0062] The inventors further believe the best method of assembly is as follows:

[0063] Step 1. Cast the porous polycarbonate film with a liquid solution made of the mixture of comb siloxane polymer, poly(ethylene glycol) ethyl ether methacrylate, poly(ethylene glycol) dimethacrylate, lithium salt $\text{Li}(\text{CF}_3\text{SO}_2)_2\text{N}$, and a radical initiator.

[0064] Step 2. Cast the porous positive electrode made of oxide active material, PVDF binder and carbon additive with the liquid solution made of the mixture of comb siloxane polymer, poly(ethylene glycol) ethyl ether methacrylate, poly(ethylene glycol) dimethacrylate, lithium salt $\text{Li}(\text{CF}_3\text{SO}_2)_2\text{N}$, and a radical initiator.

[0065] Step 3. Cast the porous negative electrode made of carbon and PVDF binder with the liquid solution made of the mixture of comb siloxane polymer, poly(ethylene glycol) ethyl ether methacrylate, poly(ethylene glycol) dimethacrylate, lithium salt $\text{Li}(\text{CF}_3\text{SO}_2)_2\text{N}$, and a radical initiator.

[0066] Step 4. Wind the cast porous polycarbon film with the cast positive and negative electrode in prismatic wound configuration.

[0067] Step 5. Put the jelly roll in a flexible packaging and seal.

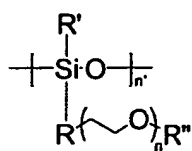
[0068] Step 6. Cure the cell at 80°C for 1 h.

[0069] Step 7. Seal the cell packaging.

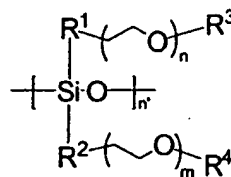
[0070] Having described the present invention, it should be apparent to the reader that
5 many variations of the present invention are possible without departure from the scope of the
present invention. The specific implementations disclosed above are by way of example and
for the purposes of enabling persons skilled in the art to implement the invention only.
Accordingly, the invention is not to be limited except by the appended claims and legal
equivalents.

Claims:

1. In an electrochemical cell, an interpenetrating network solid polymer electrolyte comprising at least one branched siloxane polymer having one or more poly(alkylene oxide) branch as a side chain, at least one crosslinking agent, at least one monofunctional monomeric compound for controlling crosslinking density, at least one metal salt and at least one radical reaction initiator.
2. The interpenetrating network solid polymer electrolyte of claim 1, wherein said poly(alkylene oxide) side chain of siloxane polymer is represented by formulas (I-a and I-b) as a metal ion conducting phase,



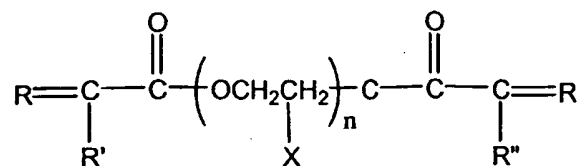
(I-a)



(I-b)

where each of R, R¹ and R² represents oxygen or a group selected from an alkylene oxide group having 1 to 6 carbon atoms; each of R', R'', R³ and R⁴ represents hydrogen or a group selected from an alkyl group having 1 to 12 carbon atoms and/or an alkenyl group having 2 to 8 carbon atoms; each of n and m represents whole numbers from 1 to 12; and n' represents whole numbers from 4 to 10,000.

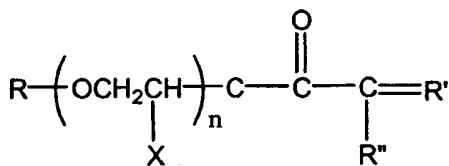
- 3 The interpenetrating network solid polymer electrolyte of claim 1, wherein said crosslinking agent is represented by formula (II),



(II)

where R represents a group selected from an alkyl group having 1 to 10 carbon atoms;
 5 and each of R' and R'' represents hydrogen or a group selected from an alkyl group
 having 1 to 10 carbon atoms and/or an alkenyl group having 2 to 12 carbon atoms; X
 is hydrogen or methyl group; and n represents a whole number from 1 to 15.

4. The interpenetrating network solid polymer electrolyte of claim 1, wherein said
 10 monomeric unit for controlling crosslinking density is represented by formula (III),



(III)

where each of R and R' represents a group selected from an alkyl group having 1 to
 10 carbon atoms; and R'' represents hydrogen or a group selected from an alkyl group
 15 having 1 to 10 carbon atoms and/or an alkenyl group having 2 to 12 carbon atoms; X
 is hydrogen or a methyl group; and n represents a whole number from 1 to 20.

5. The interpenetrating network solid polymer electrolyte of claim 2, wherein the
 branched siloxane polymer contained therein in a proportion of 10 to 80 percent by
 20 weight of total weight of organic compounds in the solid polymer electrolyte.

6. The interpenetrating network solid polymer electrolyte of claim 2, wherein said proportion of said branched siloxane polymer is 30 to 75 percent by weight of total weight of organic compounds in the solid polymer electrolyte.
- 5 7. The interpenetrating network solid polymer electrolyte of claim 2, wherein said proportion of said the branched siloxane polymer is 50 to 70 percent by weight of total weight of organic compounds in the solid polymer electrolyte.
8. The interpenetrating network solid polymer electrolyte of claim 3, wherein said
10 crosslinking agent is contained in a proportion of 5 to 60 percent by weight of total weight of organic compounds in the solid polymer electrolyte.
9. The interpenetrating network solid polymer electrolyte of claim 3, wherein said crosslinking agent is contained in a proportion of 10 to 40 percent by weight of total
15 weight of organic compounds in the solid polymer electrolyte.
10. The interpenetrating network solid polymer electrolyte of claim 4, wherein said monomeric compound exists in a proportion of 10 to 50 percent by weight of total weight of organic compounds in the solid polymer electrolyte.
- 20 11. The interpenetrating network solid polymer electrolyte of claim 4, wherein said monomeric compound exists in a proportion of 15 to 40 percent by weight of total weight of organic compounds in the solid polymer electrolyte.
- 25 12. The interpenetrating network solid polymer electrolyte of claim 1, wherein said at least one metal salt is a lithium salt.

13. An interpenetrating network polymer electrolyte of claim 12, wherein said lithium salt comprises one or more of the following: LiClO_4 , LiBF_4 , LiAsF_6 , LiPF_6 , LiCF_3SO_3 , $\text{Li}(\text{CF}_3\text{SO}_2)_2\text{N}$, $\text{Li}(\text{CF}_3\text{SO}_2)_3\text{C}$, $\text{LiN}(\text{SO}_2\text{C}_2\text{F}_5)_2$, lithium alkyl fluorophosphates, LiBOB, and a mixture thereof.
14. The interpenetrating network solid polymer electrolyte of claim 12, wherein molar ratio of said lithium salt relative to the total molar concentration of oxygen in all of the organic compounds in the polymer electrolyte is 0.01 to 0.2.
15. The interpenetrating network solid polymer electrolyte of claim 1, wherein said at least one radical reaction initiator is a thermal initiator.
16. The interpenetrating network polymer electrolyte of claim 15, wherein the thermal initiator is selected from azo compounds, peroxide compounds, bismaleimide and mixtures thereof.
17. The interpenetrating network polymer electrolyte of claim 16, wherein said azo compounds include azoisobutyronitrile.
18. The interpenetrating network polymer electrolyte of claim 16, wherein said peroxide compounds include benzoylperoxide.
19. The interpenetrating network polymer electrolyte of claim 1, wherein said electrolyte is incorporated into a porous medium.

20. The interpenetrating network polymer electrolyte of claim 19, wherein said porous medium is selected from polyolefin separator, polyolefin nonwoven type separator and polycarbonate microporous membrane.

5

21. A method for preparing the interpenetrating network polymer electrolyte of claim 1, comprising the steps of:

a) dissolving a lithium salt and a radical initiator in a branched siloxane polymer; mixing at least one crosslinking agent and a monomeric compound with the resulting solution;

10

b) casting the resulting mixture onto a substrate; and placing the cast liquid film in an oven or a heating medium such as hot plate for solidification thereof.

22. The method of claim 21 wherein said substrate is a porous medium.

15

23. The method of claim 21 wherein said substrate is a surface of an electrode.

24. A lithium ion rechargeable cell comprising of at least one lithium metal or lithium alloy anode, the solid polymer electrolyte of claim 1, and at least one metal oxide cathode.

20

25. A lithium rechargeable cell comprising at least one carbon anode, interpenetrating network solid polymer electrolyte of claim 1, and at least one metal oxide cathode.

25

26. A method for assembling a lithium rechargeable cell with the solid polymer electrolyte, comprising the steps of:

- a) coating at least one branched siloxane polymer having one or more poly(alkylene oxide) as a side chain, at least one crosslinking agent, at least one monofunctional monomeric compound for controlling crosslinking density, at least one metal salt and at least one radical reaction initiator onto one or more surfaces of a porous supporter, a cathode laminate and anode laminate;
- b) curing the precursor solution to make solid polymer electrolyte;
- c) stacking each components including porous supporter, cathode laminate and anode laminate;
- d) winding or folding the stacked components to prepare spiral wound cell or prismatic cell; and
- e) packaging the cell in a metal can, plastic pouch or foil-plastic laminated pouch.

27. A method for assembling a lithium ion rechargeable cell with the solid polymer electrolyte, comprising the steps of:

- a) coating at least one branched siloxane polymer having one or more poly(alkylene oxide) as a side chain, at least one crosslinking agent, at least one monofunctional monomeric compound for controlling crosslinking density, at least one metal salt and at least one radical reaction initiator onto one or more surfaces of a porous supporter, a cathode laminate and an anode laminate;
- b) stacking each component including said porous supporter, said cathode laminate and said anode laminate;
- c) winding or folding the stacked components to prepare spiral wound cell or prismatic cell;

d) curing the cell to change the precursor solution to solid polymer electrolyte; and packaging the cell with metal can, plastic pouch, or foil-plastic laminated pouch..

5 28. A method for assembling a lithium rechargeable cell with the solid polymer electrolyte of claim 8, comprising the steps of:

a) stacking each component including a porous supporter, at least one cathode laminate and at least one anode laminate to assemble the cell;

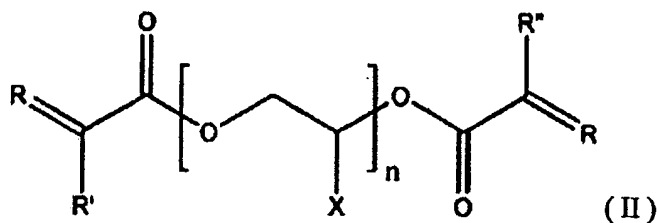
10 b) winding or folding the stacked components to prepare a spiral wound cell or prismatic cell; putting the cell in a metal can, plastic pouch or foil-plastic laminated pouch; injecting the mixture of the electrolyte components listed in claim 1 into the pre-assembled cell; and

c) curing the cell to change the precursor solution to solid polymer electrolyte.

AMENDED CLAIMS

[received by the International Bureau on 05 August 2003 (05.08.03);
original claims 1-28 replaced by new claims
1-52. (9 pages)]

1. An electrochemical device, comprising:
an electrolyte including
5 at least one siloxane polymer including one or more poly(alkylene oxide) side
chains; and
a second polymer forming an interpenetrating polymer network with the at least one
siloxane polymer.
2. The electrochemical device of claim 1, wherein at least one siloxane polymer is not
10 covalently bonded to the second polymer.
3. The electrochemical device of claim 1, wherein the at least one siloxane polymer is
not cross-linked.
4. The electrochemical device of claim 1, wherein the second polymer is cross-linked.
5. The electrochemical device of claim 1, wherein the second polymer has at least one
15 side chain including a poly(alkylene oxide) moiety.
6. The electrochemical device of claim 1, wherein the electrolyte is a solid.
7. The electrochemical device of claim 1, wherein the second polymer is selected so as
to be fabricated from a monomer represented by formula (II),



20

where

R is an alkyl group having 1 to 10 carbon atoms;

each of R' and R'' is selected from the group consisting of: hydrogen, an alkyl group
having 1 to 10 carbon atoms, and an alkenyl group having 2 to 12 carbon atoms;

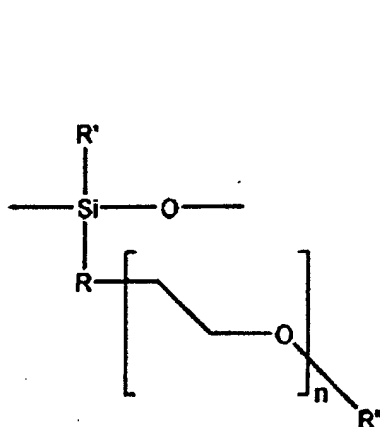
X is hydrogen or a methyl group; and

AMENDED SHEET (ARTICLE 19)

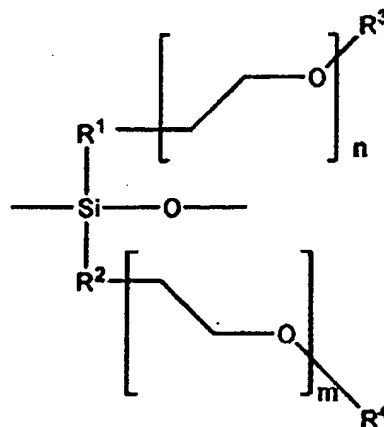
n is a whole number from 1 to 15.

8. The electrochemical device of claim 1, wherein the alkylene oxide includes ethylene oxide.

9. The electrochemical device of claim 1, wherein at least one silicon in the siloxane polymer main chain is linked to one or more side chains as represented by formula I-a or formula I-b



(I-a)



(I-b)

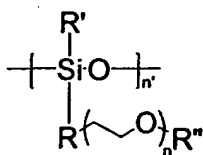
where

each of R, R¹ and R² is selected from the group consisting of: oxygen and an alkylene oxide group having 1 to 6 carbon atoms;

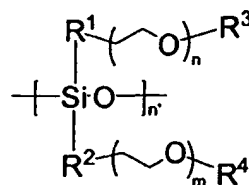
each of R', R'', R³ and R⁴ is selected from the group consisting of: hydrogen, an alkyl group having 1 to 12 carbon atoms, and an alkenyl group having 2 to 8 carbon atoms; and

each n and m is a whole number from 1 to 12.

10. The electrochemical device of claim 1, wherein the siloxane polymer is represented by a formula I-a or I-b



(I-a)



(I-b)

where

each of R, R¹ and R² is selected from the group consisting of: oxygen and an alkylene oxide group having 1 to 6 carbon atoms;

5 each of R', R'', R³ and R⁴ is selected from the group consisting of: hydrogen, an alkyl group having 1 to 12 carbon atoms, and an alkenyl group having 2 to 8 carbon atoms;

each of n and m is a whole number from 1 to 12; and

n' is a whole number from 10 to 10,000.

10 11. The electrochemical device of claim 1, wherein the proportion of the siloxane polymer is 10 to 80 weight percent of the total weight of the polymer in the electrolyte.

12. The electrochemical device of claim 1, wherein the proportion of the siloxane polymer is 50 to 70 weight percent of the total weight of the polymer in the electrolyte.

13. The electrochemical device of claim 1, wherein the second polymer is 5 to 60 weight percent of the total weight of the polymer in the electrolyte.

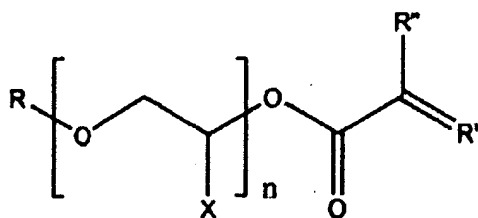
15 14. The electrochemical device of claim 1, wherein the second polymer is 10 to 40 weight percent of the total weight of the polymer in the electrolyte.

15. The electrochemical device of claim 1, wherein the electrolyte is incorporated into a porous medium.

20 16. The electrochemical device of claim 15, wherein the porous medium includes one or more components selected from the group consisting of: a polyolefin separator, a polyolefin nonwoven type separator, and a polycarbonate microporous membrane.

17. The electrochemical device of claim 1, where the electrolyte further comprises:
at least one monofunctional monomeric compound for controlling cross-linking density, at least one metal salt, and at least one initiator.

18. The electrochemical device claim 17, wherein the monomeric unit for controlling cross-linking density is represented by formula (III),



(III)

- 5 where
 each of R and R' is an alkyl group having 1 to 10 carbon atoms;
 R'' is selected from the group consisting of: hydrogen, an alkyl group having 1 to 10
 carbon atoms, and an alkenyl group having 2 to 12 carbon atoms;
 X is hydrogen or a methyl group; and
 10 n is a whole number from 1 to 20.

19. The electrochemical device of claim 17 wherein the monomeric compound is 10 to 50 weight percent of the total weight of the polymer in the electrolyte.

20. The electrochemical device of claim 17, wherein at least one metal salt is a lithium salt.

- 15 21. The electrochemical device of claim 20, wherein the lithium salt includes one or more components selected from the group consisting of LiClO₄, LiBF₄, LiAsF₆, LiPF₆, LiCF₃SO₃, Li(CF₃SO₂)₂N, Li(CF₃SO₂)₃C, LiN(SO₂C₂F₅)₂ and lithium alkyl fluorophosphates.

22. The electrochemical device of claim 20, wherein a molar ratio of the lithium salt relative to the molar concentration of oxygen in the solid polymer electrolyte is in a range of
 20 0.01 to 0.2.

23. The electrochemical device of claim 17, wherein the initiator is a thermal initiator.

24. The electrochemical device of claim 17, wherein the initiator includes one or more components selected from a group consisting: of azo compounds, peroxide compounds, and bismaleimide.

25. The electrochemical device of claim 17, wherein the initiator includes azoisobutyronitrile.

26. The electrochemical device of claim 17, wherein the initiator includes benzoylperoxide.

5 27. The electrochemical device of claim 17, wherein the second polymer includes a polyalkyldiacrylate.

28. The electrochemical device of claim 17, wherein the second polymer includes a polyalkyldimethacrylate.

10 29. The electrochemical device of claim 17, wherein the second polymer is one of a plurality of second polymers.

30. A method of forming an electrochemical device, comprising:

obtaining an electrolyte including

at least one siloxane polymer including one or more poly(alkylene oxide) side chains; and

15 a second polymer forming an interpenetrating polymer network with the at least one siloxane polymer; and

activating an anode and a cathode with the electrolyte.

31. The method of claim 30, wherein

20 obtaining the electrolyte includes mixing at least one siloxane polymer with a second polymer precursor such that the second polymer precursor forms a second polymer that forms an interpenetrating polymer network with the at least one siloxane polymer.

25 32. The method of claim 31, wherein mixing at least one siloxane polymer with a second polymer precursor includes mixing at least one compound for terminating polymerization of the second polymer, at least one metal salt, and at least one initiator.

33. The method of claim 30, wherein the interpenetrating network is formed such that at least one siloxane polymer is not covalently bonded to the second polymer.

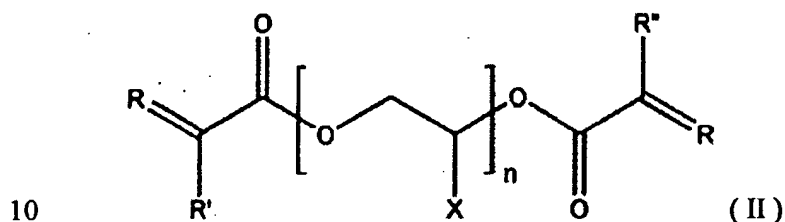
34. The method of claim 30, wherein the interpenetrating network is formed such that at least one siloxane polymer is not cross-linked.

35. The method of claim 30, wherein the interpenetrating network is formed such that the second polymer is cross-linked.

5 36. The method of claim 30, wherein obtaining the electrolyte includes forming a second polymer having one or more side chains that include a poly(alkylene oxide) moiety.

37. The method of claim 30, wherein the electrolyte is a solid.

38. The method of claim 30, wherein obtaining the electrolyte includes forming the second polymer from a monomer represented by formula (II),



where

R is an alkyl group having 1 to 10 carbon atoms;

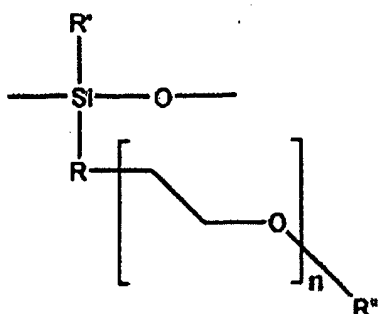
each of R' and R'' is selected from the group consisting of: hydrogen, an alkyl group having 1 to 10 carbon atoms, and an alkenyl group having 2 to 12 carbon atoms;

15 X is hydrogen or a methyl group; and

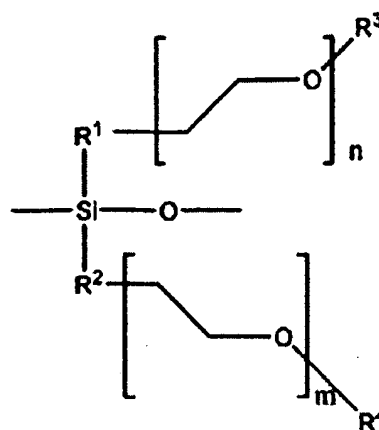
n is a whole number from 1 to 15.

39. The method of claim 30, wherein the alkylene oxide includes ethylene oxide.

40. The method of claim 30, wherein at least one silicon in the siloxane polymer main chain is linked to one or more side chains as represented by formula I-a or formula I-b



(I-a)



(I-b)

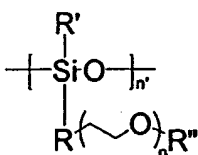
where

each of R, R¹ and R² is selected from the group consisting of: oxygen and an alkylene
oxide group having 1 to 6 carbon atoms;

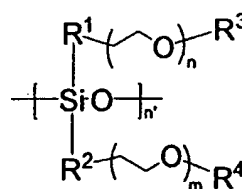
each of R', R'', R³ and R⁴ is selected from the group consisting of: hydrogen, an alkyl
group having 1 to 12 carbon atoms, and an alkenyl group having 2 to 8 carbon
atoms; and

each of n and m are from 1 to 12.

41. The method of claim 30, wherein obtaining the electrolyte includes forming a
siloxane polymer represented by a formula I-a or I-b



(I-a)



(I-b)

where

each of R, R¹ and R² is selected from the group consisting of: oxygen and an alkylene
oxide group having 1 to 6 carbon atoms;

each of R', R'', R³ and R⁴ is selected from the group consisting of: hydrogen, an alkyl
group having 1 to 12 carbon atoms, and an alkenyl group having 2 to 8 carbon
atoms;

each of n and m is a whole number from 1 to 12; and

n' is a whole number from 10 to 10,000.

42. The method of claim 30, wherein the proportion of the siloxane polymer is 10 to 80 percent by weight of the total weight of the solid polymer electrolyte.

43. The method of claim 30, wherein the proportion of the siloxane polymer is 50 to 70 percent by weight of the total weight of the solid polymer electrolyte.

44. The method of claim 30, wherein the second polymer is 10 to 40 percent by weight of the total weight of solid polymer electrolyte.

45. The method of claim 30, further comprising: incorporating the electrolyte into a porous medium before activating the anode and the cathode.

46. The method of claim 45, wherein the porous medium includes one or more components selected from the group consisting of: a polyolefin separator, a polyolefin nonwoven type separator, and a polycarbonate microporous membrane.

47. The method of claim 30, wherein obtaining the electrolyte includes forming a polyalkyldiacrylate to serve as the second polymer.

48. The method of claim 30, wherein obtaining the electrolyte includes forming a polyalkyldimethacrylate to serve as the second polymer.

49. A method of forming an electrochemical cell, comprising:
coating one or more surfaces of a porous supporter, a cathode laminate, and anode laminate with a precursor solution, the precursor solution including a siloxane polymer having one or more side chains including a poly(alkylene oxide) moiety, at least one crosslinking agent, at least one monofunctional monomeric compound for controlling crosslinking density, at least one metal salt, and at least one radical reaction initiator;
curing the precursor solution to make a solid electrolyte; and
stacking the porous supporter, cathode laminate, and anode laminate.

50. The method of claim 49, further comprising:

winding or folding the stacked components to form a cell.

51. The method of claim 49, wherein stacking the porous supporter, cathode laminate, and anode laminate is performed before curing the precursor solution to make a solid electrolyte.

- 5 52. The method of claim 49, further comprising:
packaging the stacked porous supporter, cathode laminate and anode laminate in a metal
can or pouch; and
injecting the precursor solution into the metal can or pouch before curing the precursor
solution to make the solid electrolyte.

10

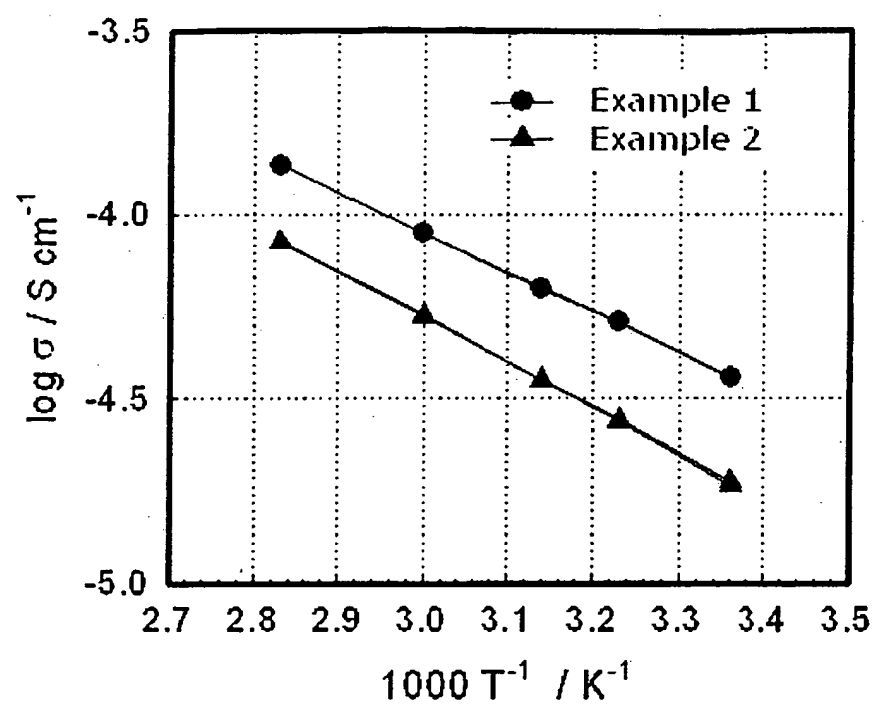
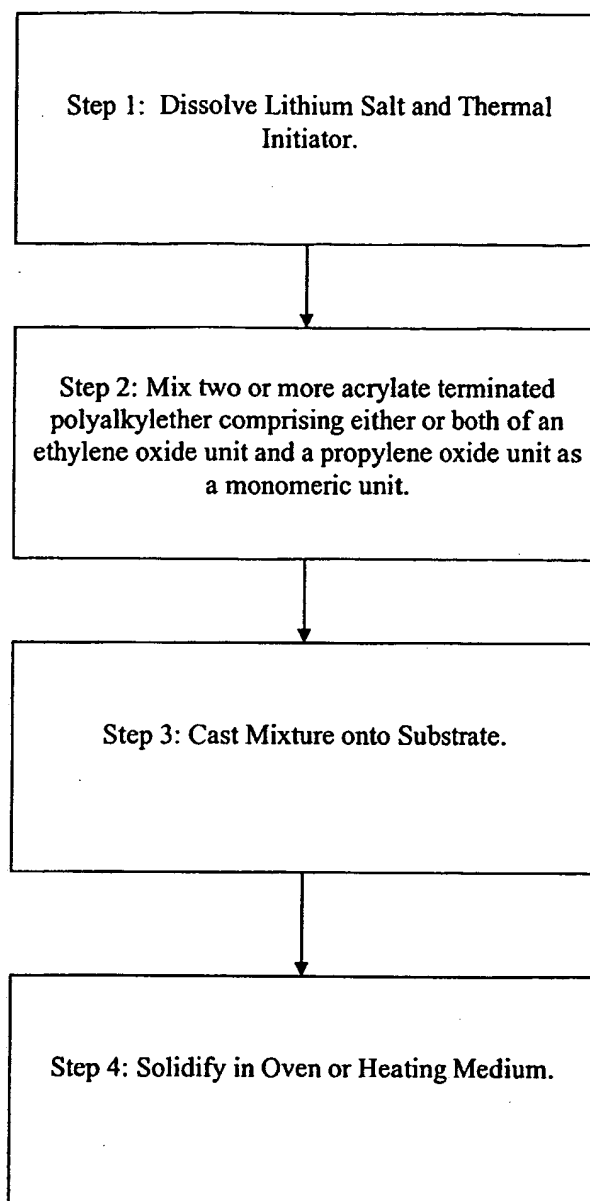


Figure 1

**FIGURE 2**

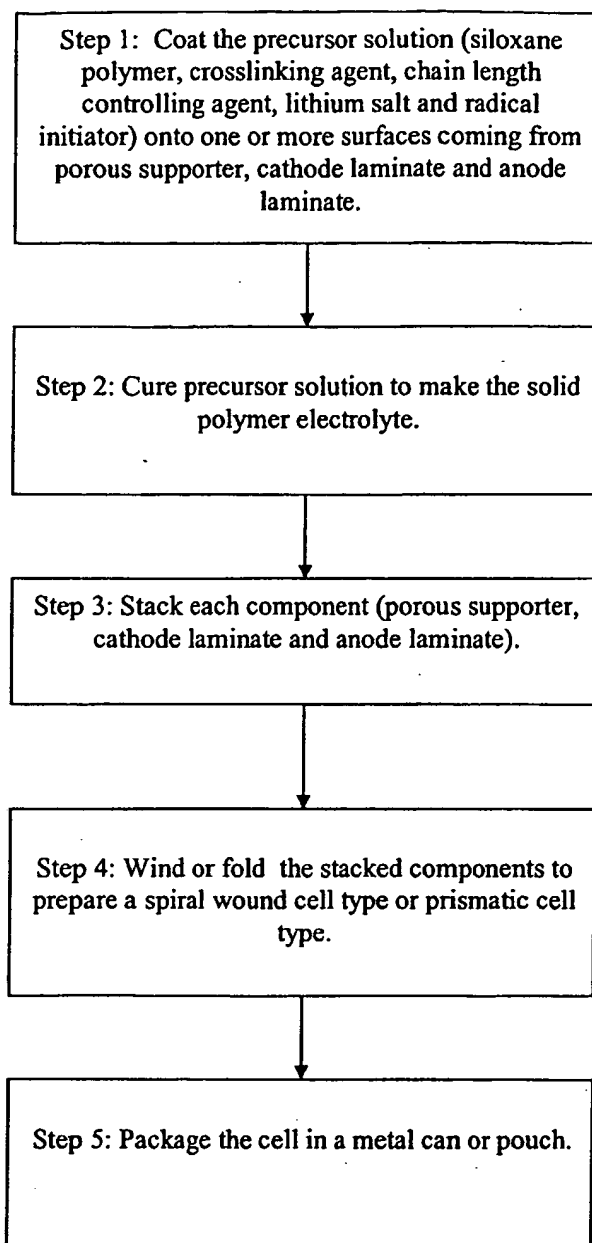


Figure 3

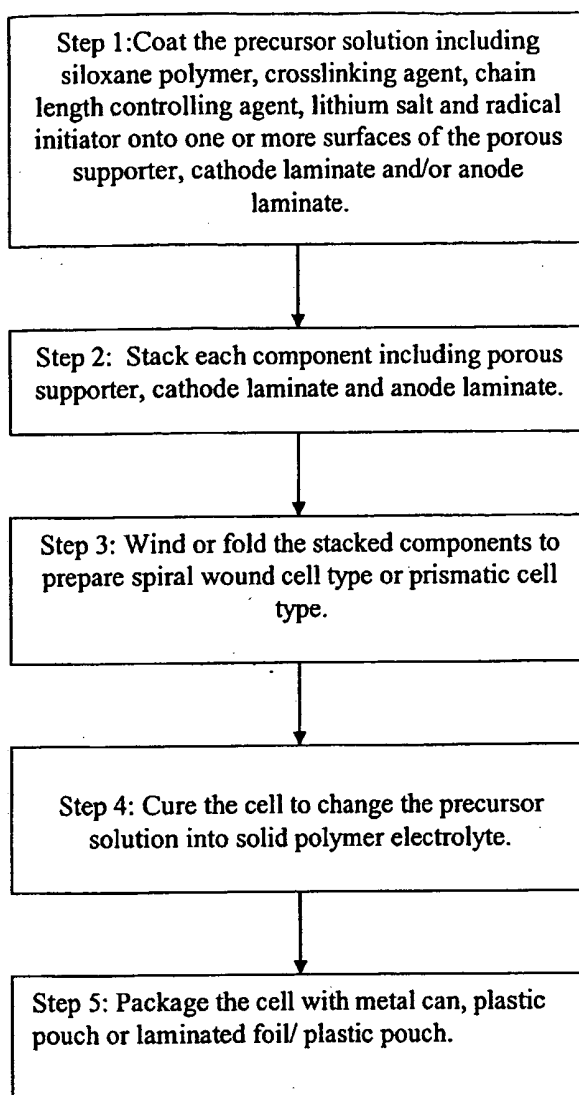


FIGURE 4

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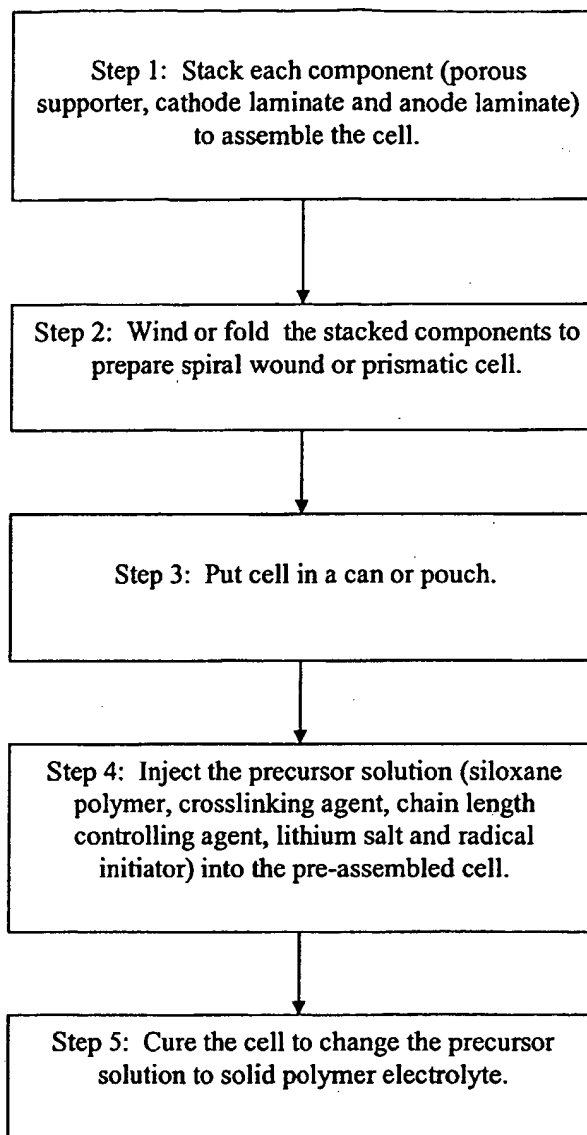


FIGURE 5

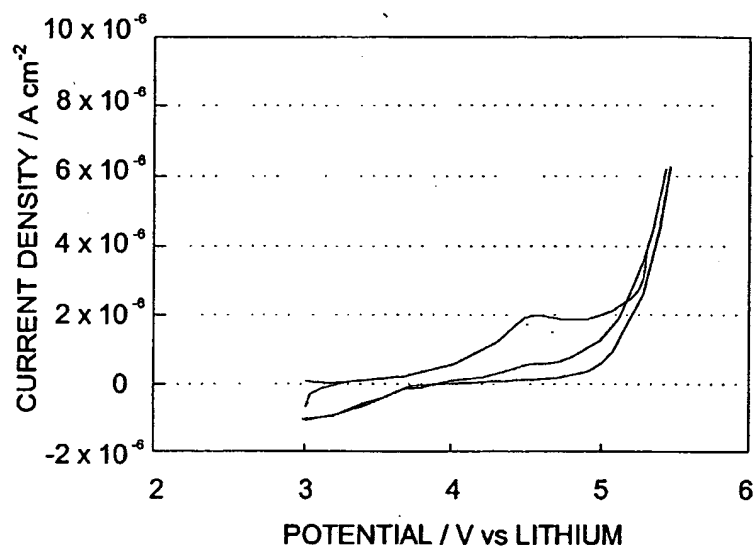


FIGURE 6

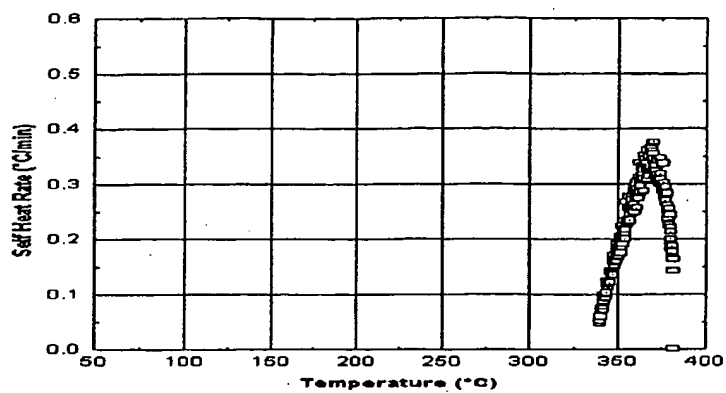


Figure 7a (Comparison 1)

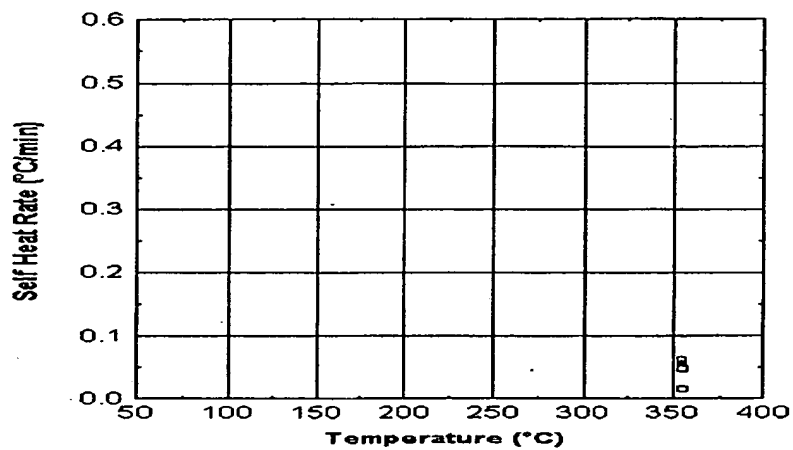


Figure 7b (Comparison 2)

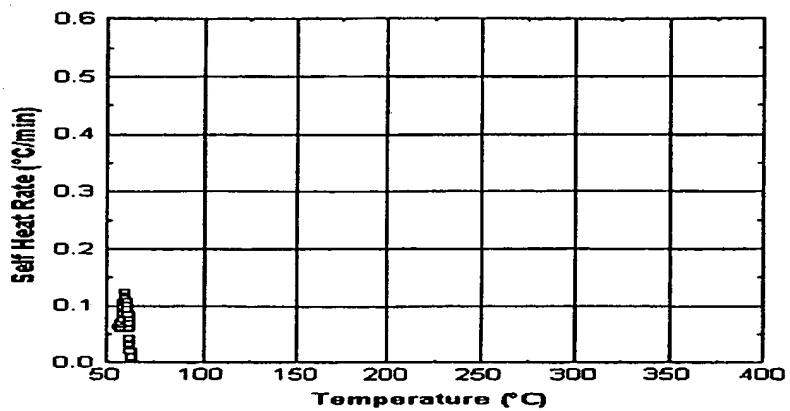


Figure 7c (Example 4)

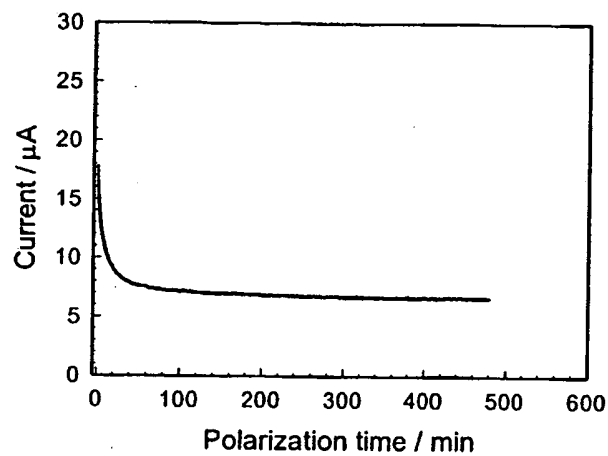


Figure 8a

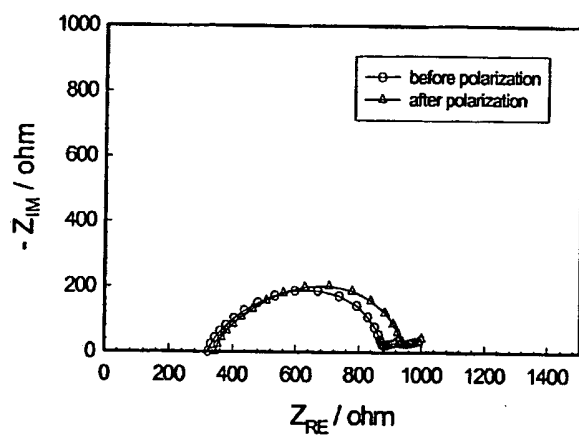


Figure 8b

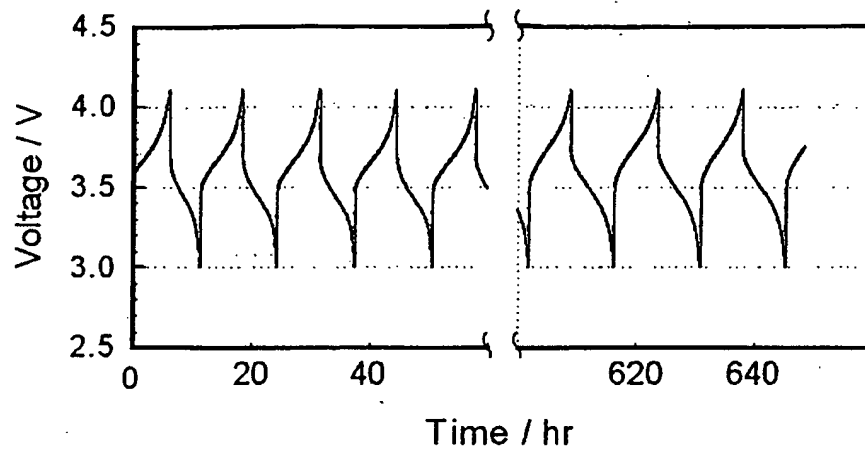


Figure 9a

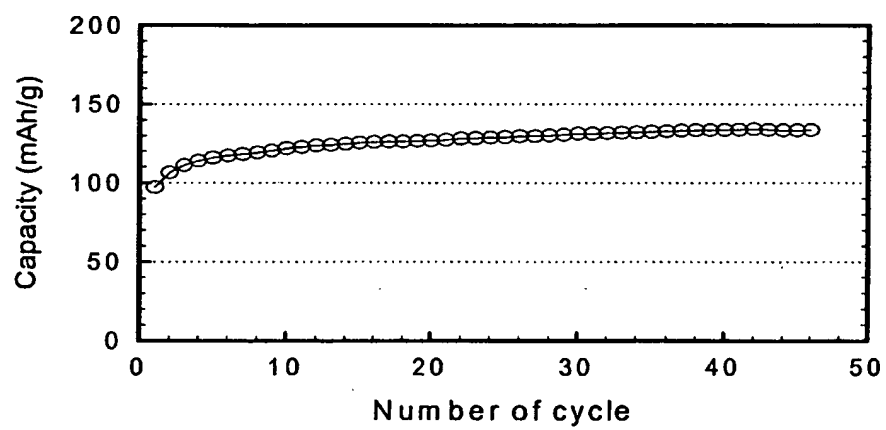


Figure 9b

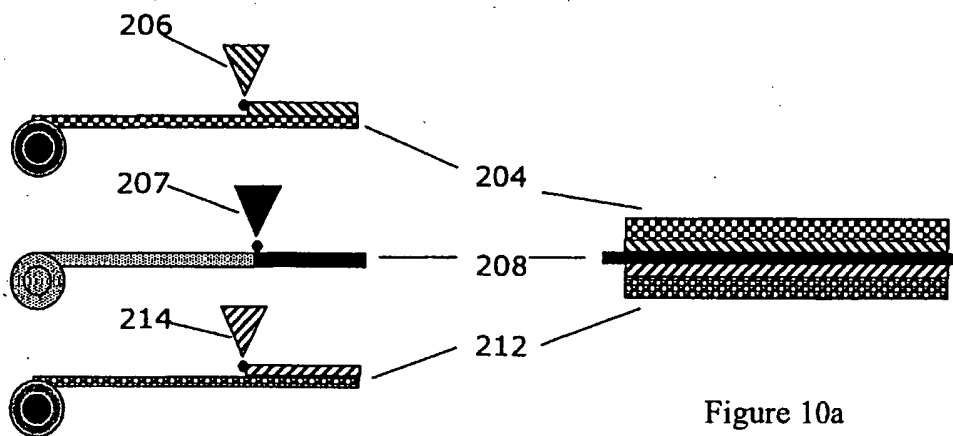


Figure 10a

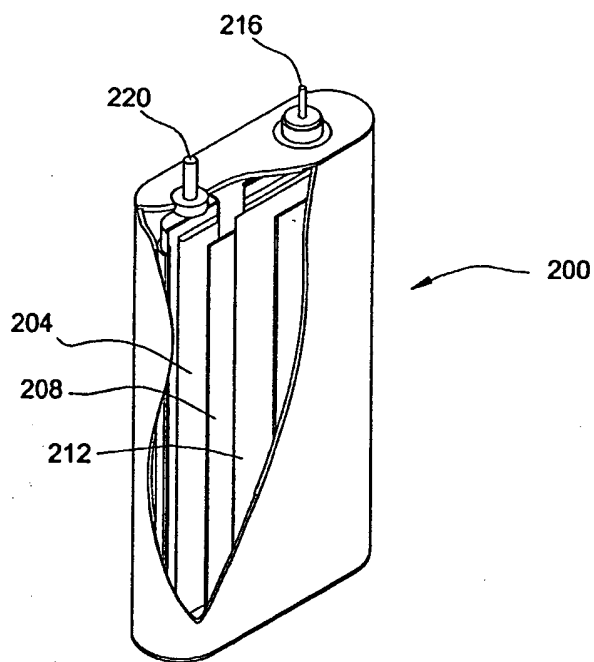


FIGURE 10b

INTERNATIONAL SEARCH REPORT

International application No.

PCT/US03/02128

A. CLASSIFICATION OF SUBJECT MATTER

IPC(7) : H01M 6/14, 6/18, 10/40; H01G 9/025

US CL : 252/62.2; 429/188, 302-304, 313, 321-323; 361/504, 505, 523

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

U.S. : 252/62.2; 429/188, 302-304, 313, 321-323; 361/504, 505, 523

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
Y	US 5,112,512 A (NAKAMURA) 12 May 1992 (12.05.1992), see abstract and col. 3, lines 23-col. 5, lines 15.	1-28
Y	US 5,419,984 A (CHALONER-GILL et al) 30 May 1995 (30.05.1995), col. 4, lines 22-63 and col. 5, lines 18-col. 6, lines 53.	1-28
Y	US 5,885,733 A (OHSAWA et al) 23 March 1999 (23.03.1999), see col. 2, lines 25-62 and col. 4, lines 19-67.	1-28
A	US 5,300,375 A (CHALONER-GILL) 05 April 1994 (05.04.1994).	1-28
A	US 5,362,493 A (SKOTHEIM et al) 08 November 1994 (08.11.1994).	1-28
A	US 5,538,812 A (LEE et al) 23 July 1996 (23.07.1996).	1-28
A	US 6,124,062 A (HORIE et al) 26 September 2000 (26.09.2000).	1-28
A	US 5,690,702 A (SKOTHEIM et al) 25 November 1997 (25.11.1997).	1-28

☐ Further documents are listed in the continuation of Box C.

☐ See patent family annex.

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Date of the actual completion of the international search

12 June 2003 (12.06.2003)

Date of mailing of the international search report

11 JUL 2003

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